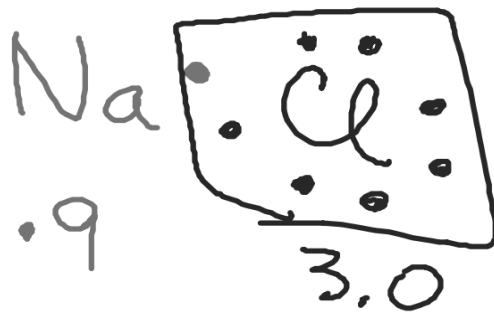
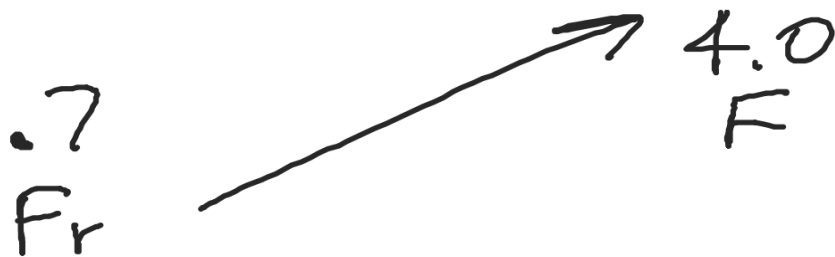


IONIC BOND -  
TRANSFER of electron  
to the stronger Atom

NaCl



ELECTRONEGATIVITY -  
The ability of an atom  
TO ATTRACT AN  $e^-$



electronegativity Difference  
end

high - lowest en

$$\begin{array}{ccc} 3.0 & - & .9 \\ \text{Cl} & & \text{Na} \end{array} = 2.1$$

COVALENT BOND  
the electrons are shared

POLAR - A medium end  
end .4 - 2.0

NON-POLAR - A low end  
end < .4

ALWAYS GO UP if @ 2.0 or  
-4

#2  $H_2O$



$$H = 2.1$$

$$O = 3.5$$

$$3.5 - 2.1$$

1.4 polar

BOND IS BETWEEN 2<sup>covalent</sup>  
elements .4 1.4 2.0

